S.S.R. DEGREE COLLEGE, (AUTONOMOUS)
NIZAMABAD (C.C:5029)
I SEMESTER INTERNAL ASSESSMENT I EXAMINATIONS
CHEMISTRY
QUESTION BANK

I. Multiple Choice Questions

1. The cations of group III are _____ (A)
   a) Co^{2+}, NI^{2+}.Zn^{2+}, Mn^{2+} b) Cu^{2+}, Cd^{2+}, Pb^{2+}
   C) Hg^{2+}, Pb^{2+} , BI^{3+} d) None
2. The ion that can't be precipitated by both HCL and H_2s is. (B)
   a) Cu^{+} b) Sn^{2+} c) Pb^{2+} d) Ag^{2+}
3. Which one of the following has minimum solubility product (D)
   a) $Alcl_3$ b) NH_4Cl c) $Bacl_2$ d) Agcl
4. Potassium cyanide is used to separate the following ions (B &C)
   a) Mn^{2+} and Zn^{2+} b) Cu^{2+} and Cd^{2+}
   c) Co^{2+} and Ni^{2+} d) Ba^{2+} and Ca^{2+}
5. _____ ions give crimsom colour in flame (A)
   a) Sr^{2+} b) Ca^{2+} c) K^{+} d) Ba^{2+}
6. _____ gives a white precipitate with aqueous $AgNo_3$ and a green flame test (C)
   a) Kcl b) $CaCl_2$ c) $Bacl_2$ d) Nacl
7. Which of the following sulphates is insoluable in water? (D)
   a) $Cuso_4$ b) $Cdso_4$ c) $Bi(so_4)_3$ d) $Pbso_4$
8. _____ions cannot be separated by H s in dilute hydrochloric acid (D)
   a) Al^{3+} , Hg^{2+} b) Ni^{2+} , Cu^{2} c) Zn^{2+} , Cu^{2+} d) Bi^{3+} , Sn^{4+}
9. Which of the following cation will not be precipitated by H2s in presence of ammonia? (A)
   a) Cd^{2+} b) fe^{3+} c) Mn^{2+} d) Co^{2+}
10. The group reagent of group V radicals is_____ (A)
    a) $(NH_4)_2Co_3$ in presence of NH_4OH b) H_2s
    c) HCL d) None
11. An example of basic oxide is____ (A)
    a) Na_2O b) N_2O_5 c)Co_2 d) Co
12. Superoxides are in nature (A)
    a) Paramagnetic b) Diamagnetic c) Both (a)&(b) d) Neutral
13. Pb_3O_4 is a _____ (C)
    a) Suboxide b) Neutrol oxide
    c) Mixed oxide d) Amphoteric oxide.

14. Carbon monoxide is a resonance hydrid of how many structures? (B)
    a) 2 b) 3 c) 4 d) None
15. _____ are the most volatile constituents. (B)
    a) Kr and Xe b) He and Ne c) Ar and Kr d) Ne and Xe
16. The structure of $XeOf_4$ is _____ (C)
    a) Square pyramidal b) Square bipyramidal
    c) Trigonal planar d) Trigonal bipyramidal
17. Separation of gases from one another is carried out by a method called.. (A)
    a) Dewar's method b) Rayleigh's method
    c) Physical method d) None of the above
18. Xef_4 when treated with Bcl_3, gives_____ (B)
    a) Xef_2 b) Bf_3 c)HF d) None of the above
19. The formula of nitrous acid is _____ (A)
    a) HNO_2 b) HNO_3 c) HNO_4 d) HOONO
20. The first three elements of _____ groups are called as chromium triad. (C)
    a) Fourth b) Fifth c) Sixth d) Eleventh

II. FILL IN THE BLANKS

1. The atomic radius of group 13 elements, **Increases** from boron to thallium.
2. The hydrolysis of diborane produces . **Bovie acid** and **hydrogen**.
3. The structure of **Borazine** resembles with that of benzene.
4. **Silicones** are the organosilion polymers. Having. Si– O-si linkages.
5. **Decomposition** of hydroxlylamine produces nitrogen, ammonia and water.
6. The process of precipitation and dissolution are carried out based on the priniciple of **solubility product**.
7. All the carbonates decompose when react with diluted . **HCL (or) H_2SO_4**.
8. **Blue** precipitate is obtained by adding potassium ferricyanide to a solution containing Fe^{2+} ions.
9. The colour of zine sulphide is **White**
10. **Ammonium thiocyanate** is used in the detection of CO^{2+} and Fe^{3+} ions.
11. Acidic oxides are formed when oxygen reacts with Non metals.
12. The electronic configuration of xenon in XeO_3 is $5s^2\ 5p^3\ 5d^3$.
13. Pseudohalogens are also called as Halogenoids.
14. β –quinol is a clatherate compound.
15. When a mixture of noble gases are brought into contact with charcoal at -100^0c, the gases absorbed at charcoal are Argon, loxypton & xenon.
16. The bond angle in Xef_2 is **180**
17. The geometry of XeO_3 is **Trigonal planar.**
18. In laboratories XeO_3 is used to measure alcohols, carboxylic acids & **Amino Acids**
19. **Covalent** & **Hydrogen** types of bonds are formed in boric acid.
20. The thermodynamic stability of the compounds increases with **Decrease** in ionization energy.

III. Descriptive Questions

1. Interhalogen
2. Definition & types of oxides
3. Markovnikov's & Anti markovnikov's rule
4. Orientation
5. Nitration.