

TELANGANA UNIVERSITY
S.S.R. DEGREE COLLEGE, NIZAMABAD (C.C:5029)
II SEMESTER INTERNAL ASSESSMENT I EXAMINATIONS
CHEMISTRY QUESTION BANK

I. Multiple choice questions.

1. HNO_2 is anhydrate of [d]
a. N_2O b. NO c. NO_2 d. N_2O_3
2. How many p-o-p bonds present in P_4O_{10} [b]
a. 8 b. 6 c. 7 d. 5
3. Shape of the phosphate ion PO_4^{-3} [b]
a. Linear b. Tetrahedral c. Trigonal planar d. Octahedral
4. In IF_7 Hybridization [c]
a. SP^3d b. SP^3d^2 c. SP^3d^3 d. SP
5. Structure of ClF_5 molecule? [d]
a. Square pyramidal b. Tetra hydral c. Linear d. Trigonal planar
6. In XeF_6 undergoes hybridization [b]
a. SP^3d^2 b. SP^3d^3 c. SP^2 d. SP^3
7. Shape of XeO_3 is [a]
a. Pyramidal b. Linear c. Tetrahedral d. Trigonal planer
8. Example for AX_3 type of inter halogen is [c]
a. IF_5 b. IF_7 c. ICl_3 d. ICl
9. Shape of AX_3 type [a]
a. T.Shaped b. Linear c. Octahedral d. None
10. Shape of IF_7 mole is [b]
a. Linear b. Pentagonal bipyramidal c. Tetrahedral d. None
11. The general electronic configuration [b]
a. $(n-1)d^{1-10}ns^{1-2}$ b. $(n-1)d^{(1-10)}ns^{0-2}$ c. $(n-1)d^{10}ns^1$ d. $(n-1)d^{10}ns^1$
12. d-blocks elements are [a]
a. metals b. Non metals c. both d. None
13. Transition metals are [b]
a. Exhibit inert pair effect b. Exhibit variable oxidation states
c. Low M.P d. Don't show catalytic activity
14. Which ion gives coloured solution [d]
a. Ti^{+4} b. V^{+3} c. Cu^+ d. Zn^{+2}
15. Para magnetism is property of [a]
a. Unpaired electrons b. Paired electrons c. Transition metals d. None

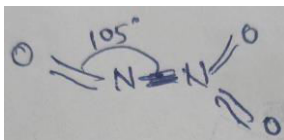
16. Diamagnetism is the property of [a]
 a. Completely filled electronic sub shells b. Unpaired electrons
 c. Transition elements d. Non-transition elements
17. The metal having the least magnetic moment value is [d]
 a. Fe^{+2} b. CO^{+2} c. Ni^{+2} d. Cu^{+2}
18. Which noble gas is filled in fluorescent lamps [c]
 a. He b. Ne c. Ar d. Xe
19. Electronic configuration of zero group elements [d]
 a. ns^2 b. np^6 c. $nd(1-10)$ d. $ns^2 np^6$
20. Which oxide of phosphorus acts as a drying agent [a]
 a. P_4O_{10} b. P_4O_6 c. P_4O_7 d. P_4O_5

II. Fill in the blanks

- An hydride of H_3PO_4 is P_4O_{10}
- Hydride of HNO_3 is N_2O_5
- Shape of CO_2 molecule is Linear
- $\text{H}_2\text{S}_2\text{O}_8$ is known as marsh's acid
- SO_3 shape is Trigonal planar
- XeOF_2 hybridization is SP^3d
- XO_4 shape is Tetrahedral
- The shape of I_3^- ion is Linear
- Example for polyhalides is ICl_2^- , ICl_4^+
- The known xenon fluorides are XeF_2 , XeF_4 and XeF_6
- Oxidation state of Mn(II) is +2
- The first xenon compound was XePtF_6
- $\text{NO} + \text{NO}_2 \rightarrow$ N_2O_3
- Fuming HNO_3 is used as an oxidizer in Rocket fuels
- Transition elements are less electropositive than the f-block element.
- Among transition elements osmium has the maximum density.
- The noble gas atoms are trapped in quinal hydrogen bond case formed compounds are called Clathrates
- Cu Traid Cu, Ag & Au
- CuCl_2 is used as catalyst in the production of chlorine by deacon's process.

III. Short Answers.

- Write example for peroxides?
A: Na_2O_2 , BaO_2
- Write example for super oxides?
A: KO_2
- Write example for mixed oxide?
A: Pb_3O_4 , Fe_3O_4
- Write the structure of N_2O_3 ?



A:

- Write phosphorus oxy acids?
A: H_3PO_3 , H_3PO_2 , $\text{H}_4\text{P}_2\text{O}_5$, $\text{H}_4\text{P}_2\text{O}_6$, $\text{H}_4\text{P}_2\text{O}_7$

6. Write oxy acids of chlorine?

A: HOCl, HClO₂, HClO₃, HClO₄

7. Write example for AX₃ molecule?

A: ICl₃, ClF₃, BrF₃

8. Write equation for magnetic moment?

A: $\mu_s = \sqrt{n(n+2)}$

9. Write Titanium Triad?

A: Ti, Zr and Hf

10. Write Cr triad?

A: Cr, Mo & W